1) For the reaction ____C_3H_8 + ____O_2 → _____CO_2 + _____H_2O

       a. How many moles of CO_2 will be produced from 2.53 mole of C_3H_8 and 10.3 mol of O_2?

       b. Which reactant is the limiting reactant? _____ Which is the excess reactant? _____

       c. If 4.35 moles of CO_2 are actually produced, what is the percent yield?

2) For the reaction: ____N_2 + ____H_2 → ____NH_3

      a. How many moles of NH_3 will be produced from 10.25 L of H_2 and 2.50 L of N_2?

       b. Which reactant is the limiting reactant? _____ Which is the excess reactant? _____

       c. If 0.297 moles of NH_3 are actually produced, what is the percent yield?

3) For the reaction ____Fe_3O_4 + ____H_2 → ____Fe + ____H_2O

    a. How many grams of Fe will be produced from 25.0 grams of Fe_3O_4 and 10.0 grams of H_2?

    b. Which reactant is the limiting reactant? _____ Which is the excess reactant? _____

    c. If 17.74 grams of Fe are actually produced, what is the percent yield?